

### 18.4 Acid-Base Chemistry: Theory

#### Multiple-Choice Questions

1. Which combination of quantities are the characteristic equations for a weak acid? \_\_\_\_\_
- (A)  $\text{pH} = \text{pK}_a + \log \frac{[\text{A}^-]}{[\text{HA}]}$       (B)  $\text{pH} = \text{pK}_a + \log \frac{[\text{HA}]}{[\text{A}^-]}$   
 (C)  $\text{pH} = \text{pK}_a + \log \frac{[\text{HA}]}{[\text{A}^-]}$       (D)  $\text{pH} = \text{pK}_a + \log \frac{[\text{A}^-]}{[\text{HA}]}$

2. Which of the following is NOT a characteristic of a weak acid? \_\_\_\_\_
- (A) It is a weak electrolyte.      (B) It is a strong electrolyte.  
 (C) It is a weak acid.      (D) It is a strong acid.

3. In a 0.1 M solution of a weak acid, the concentration of the acid is 0.05 M. What is the pH of the solution? \_\_\_\_\_
- (A) 4.0      (B) 5.0      (C) 6.0      (D) 7.0

4. In a 0.1 M solution of a weak acid, the concentration of the acid is 0.05 M. What is the pH of the solution? \_\_\_\_\_
- (A) 4.0      (B) 5.0      (C) 6.0      (D) 7.0

5. Which of the following is NOT a characteristic of a weak acid? \_\_\_\_\_
- (A) It is a weak electrolyte.      (B) It is a strong electrolyte.  
 (C) It is a weak acid.      (D) It is a strong acid.

6. In a 0.1 M solution of a weak acid, the concentration of the acid is 0.05 M. What is the pH of the solution? \_\_\_\_\_
- (A) 4.0      (B) 5.0      (C) 6.0      (D) 7.0

7. In a 0.1 M solution of a weak acid, the concentration of the acid is 0.05 M. What is the pH of the solution? \_\_\_\_\_
- (A) 4.0      (B) 5.0      (C) 6.0      (D) 7.0

8. In a 0.1 M solution of a weak acid, the concentration of the acid is 0.05 M. What is the pH of the solution? \_\_\_\_\_
- (A) 4.0      (B) 5.0      (C) 6.0      (D) 7.0



# Multiple Choice Stoichiometry Test

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