

18.4 Acid-Base Chemistry: Theory

Multiple-Choice Questions

1. Which combination of quantities are dimensional quantities? (correct)
- (A) pressure, concentration, pH, amount of substance
 (B) volume, molarity, pH, amount of substance
 (C) volume, molarity, pH, concentration of ions
 (D) amount of substance, pH, concentration of ions, concentration of substance

(A) 10^{-14} (B) 10^{-7}

(A) 10^{-14} (B) 10^{-7} (C) 10^{-14} (D) 10^{-7} (E) 10^{-14} (F) 10^{-7} (G) 10^{-14} (H) 10^{-7} (I) 10^{-14} (J) 10^{-7}

2. In a 0.1 M acetic acid solution, the concentration of acetate ions is approximately 0.01 M. What is the pH of this solution? (correct)

(A) 1.0 (B) 2.0 (C) 3.0 (D) 4.0 (E) 5.0

(A) 1.0 (B) 2.0 (C) 3.0 (D) 4.0 (E) 5.0 (F) 6.0 (G) 7.0 (H) 8.0 (I) 9.0 (J) 10.0

3. The pH of a 0.1 M acetic acid solution is approximately 3.0. What is the concentration of acetate ions in this solution? (correct)

(A) 0.1 M (B) 0.01 M (C) 0.001 M (D) 0.0001 M (E) 0.00001 M

4. Which of the following is a weak acid? (correct)



(A) HCl (B) H_2SO_4 (C) H_2CO_3 (D) HNO_3 (E) H_2O (F) H_2O_2 (G) H_2O_2 (H) H_2O_2 (I) H_2O_2 (J) H_2O_2

5. A 0.1 M solution of acetic acid has a pH of approximately 3.0. What is the concentration of acetic acid in this solution? (correct)



(A) 10^{-14} (B) 10^{-7} (C) 10^{-14} (D) 10^{-7} (E) 10^{-14} (F) 10^{-7} (G) 10^{-14} (H) 10^{-7} (I) 10^{-14} (J) 10^{-7}

6. In a 0.1 M acetic acid solution, the concentration of acetate ions is approximately 0.01 M. What is the pH of this solution? (correct)



Multiple Choice Stoichiometry Test

Padhraic Smyth



Multiple Choice Stoichiometry Test:

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