

	Reaction (Oxidised form + ne ⁻)	→ Reduced form)	E ^o / V
↑ Increasing strength of oxidising agent	F ₂ (g) + 2e ⁻	→ 2F ⁻	2.87
	Co ³⁺ + e ⁻	→ Co ²⁺	1.81
	H ₂ O ₂ + 2H ⁺ + 2e ⁻	→ 2H ₂ O	1.78
	MnO ₄ ⁻ + 8H ⁺ + 5e ⁻	→ Mn ²⁺ + 4H ₂ O	1.51
	Au ³⁺ + 3e ⁻	→ Au(s)	1.40
	Cl ₂ (g) + 2e ⁻	→ 2Cl ⁻	1.36
	Cr ₂ O ₇ ²⁻ + 14H ⁺ + 6e ⁻	→ 2Cr ³⁺ + 7H ₂ O	1.33
	O ₂ (g) + 4H ⁺ + 4e ⁻	→ 2H ₂ O	1.23
	MnO ₂ (s) + 4H ⁺ + 2e ⁻	→ Mn ²⁺ + 2H ₂ O	1.23
	Br ₂ + 2e ⁻	→ 2Br ⁻	1.09
	NO ₃ ⁻ + 4H ⁺ + 3e ⁻	→ NO(g) + 2H ₂ O	0.97
	2Hg ²⁺ + 2e ⁻	→ Hg ₂ ²⁺	0.92
	Ag ⁺ + e ⁻	→ Ag(s)	0.80
	Fe ³⁺ + e ⁻	→ Fe ²⁺	0.77
	O ₂ (g) + 2H ⁺ + 2e ⁻	→ H ₂ O ₂	0.68
	I ₂ + 2e ⁻	→ 2I ⁻	0.54
	Cu ⁺ + e ⁻	→ Cu(s)	0.52
	Cu ²⁺ + 2e ⁻	→ Cu(s)	0.34
	AgCl(s) + e ⁻	→ Ag(s) + Cl ⁻	0.22
	AgBr(s) + e ⁻	→ Ag(s) + Br ⁻	0.10
	2H ⁺ + 2e ⁻	→ H ₂ (g)	0.00
	Pb ²⁺ + 2e ⁻	→ Pb(s)	-0.13
	Sn ²⁺ + 2e ⁻	→ Sn(s)	-0.14
	Ni ²⁺ + 2e ⁻	→ Ni(s)	-0.25
	Fe ²⁺ + 2e ⁻	→ Fe(s)	-0.44
	Cr ³⁺ + 3e ⁻	→ Cr(s)	-0.74
Zn ²⁺ + 2e ⁻	→ Zn(s)	-0.76	
2H ₂ O + 2e ⁻	→ H ₂ (g) + 2OH ⁻ (aq)	-0.83	
Al ³⁺ + 3e ⁻	→ Al(s)	-1.66	
Mg ²⁺ + 2e ⁻	→ Mg(s)	-2.36	
Na ⁺ + e ⁻	→ Na(s)	-2.71	
Ca ²⁺ + 2e ⁻	→ Ca(s)	-2.87	
K ⁺ + e ⁻	→ K(s)	-2.93	
Li ⁺ + e ⁻	→ Li(s)	-3.05	

↓
Increasing strength of reducing agent

Electrode Potentials

Emilie Sanchez



Electrode Potentials:

The book delves into Electrode Potentials. Electrode Potentials is an essential topic that needs to be grasped by everyone, from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Electrode Potentials, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:

- Chapter 1: Introduction to Electrode Potentials
- Chapter 2: Essential Elements of Electrode Potentials
- Chapter 3: Electrode Potentials in Everyday Life
- Chapter 4: Electrode Potentials in Specific Contexts
- Chapter 5: Conclusion

2. In chapter 1, the author will provide an overview of Electrode Potentials. The first chapter will explore what Electrode Potentials is, why Electrode Potentials is vital, and how to effectively learn about Electrode Potentials.

3. In chapter 2, this book will delve into the foundational concepts of Electrode Potentials. This chapter will elucidate the essential principles that must be understood to grasp Electrode Potentials in its entirety.

4. In chapter 3, this book will examine the practical applications of Electrode Potentials in daily life. The third chapter will showcase real-world examples of how Electrode Potentials can be effectively utilized in everyday scenarios.

5. In chapter 4, the author will scrutinize the relevance of Electrode Potentials in specific contexts. This chapter will explore how Electrode Potentials is applied in specialized fields, such as education, business, and technology.

6. In chapter 5, the author will draw a conclusion about Electrode Potentials. The final chapter will summarize the key points that have been discussed throughout the book.

The book is crafted in an easy-to-understand language and is complemented by engaging illustrations. This book is highly recommended for anyone seeking to gain a comprehensive understanding of Electrode Potentials.

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