

# Voltaic Cells

Label the parts for the following cell:



By connecting to the voltmeter, describe a spontaneous redox reaction. (10 marks)

- a. metals \_\_\_\_\_
- b. metal ions \_\_\_\_\_
- c. anodes \_\_\_\_\_
- d. cathodes \_\_\_\_\_
- e. half-reaction oxidation: Zn \_\_\_\_\_
- f. half-reaction reduction: Ag \_\_\_\_\_

- g. write overall balanced equation \_\_\_\_\_
- h. E<sub>cell</sub> \_\_\_\_\_

i. Write overall cell process:  $\text{Zn} + 2\text{Ag}^+ \rightarrow \text{Zn}^{2+} + 2\text{Ag}$  (do not forget to balance electrons, explain the number of electrons, include half-reactions for anodes & cathodes)

\_\_\_\_\_

# Voltaic Cells Study Guide

**Joan I. Senyk, Darrell D. Ebbing**



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