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web the balanced chemical equation for this reaction is  $\text{mg} + \text{s} + 2\text{hno}_3 \text{aq} \rightarrow \text{mg} + \text{no}_3 + 2 \text{aq} + \text{h}_2 \text{g}$  in this equation the s symbolizes solid magnesium aq represents aqueous nitric acid aq signifies magnesium nitrate in a dissolved state and g represents the gaseous hydrogen produced

[\$\text{hno}\_3 + \text{mg} + \text{c}\_2\text{h}\_3\text{o}\_2 + 2 \text{hc}\_2\text{h}\_3\text{o}\_2 \rightarrow \text{mg} + \text{no}\_3 + 2 \text{chemical equation}\$](#)  - Aug 11 2023

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web mar 4 2018  $\text{mg} + 2\text{hno}_3 \rightarrow \text{mg} + \text{no}_3 + 2 \text{h}_2$  on the left side there are mg 1 magnesium in elemental form  $\text{h}_2$  there are 2 molecules of nitric acid each acid has one hydrogen so two molecules of nitric acid is 2 hydrogens n 2 there are 2 molecules of nitric acid each acid has one nitrogen so two molecules of nitric acid is 2 nitrogens o 6

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magnesium hydroxide express your answer as a chemical equation including phases 2 enter a net ionic equation for the reaction between nitric acid and magnesium hydroxide express your answer as a chemical equation including  $\text{MgCO}_3$   $\text{HNO}_3$   $\text{Mg(NO}_3)_2$   $\text{H}_2\text{O}$   $\text{CO}_2$  chemical equation - Jul 10 2023

web magnesium carbonate nitric acid magnesium nitrate water carbon dioxide one mole of magnesium carbonate  $\text{MgCO}_3$  and two moles of nitric acid  $\text{HNO}_3$  react to form one mole of magnesium nitrate  $\text{Mg(NO}_3)_2$  one mole of water  $\text{H}_2\text{O}$  and one mole of carbon dioxide  $\text{CO}_2$  show chemical structure image reactants

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*$\text{Mg(NO}_3)_2$   $\text{H}_2\text{SO}_4$   $\text{MgSO}_4$   $\text{HNO}_3$  chemical equation* - Apr 07 2023

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web how to balance  $\text{HNO}_3$   $\text{Mg}$   $\text{Mg(NO}_3)_2$   $\text{H}_2$  word equation nitric acid magnesium magnesium nitrate hydrogen gas type of chemical reaction for this reaction we have a single displacement reaction balancing strategies in this single displacement reaction the  $\text{Mg}$  replaces the h in the  $\text{HNO}_3$

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each compound to balance mg hno3 mg no3 2 no2 h2o you ll need to be

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