

Voltaic Cells

Label the parts for the following cell:



By connecting to the voltmeter, describe a spontaneous redox reaction. (10 marks)

- a. metals _____
- b. metal ions _____
- c. metals _____
- d. metal ions _____
- e. electron movement flow _____
- f. spontaneous redox reactions _____

- g. half-cell reduction half-cell _____
- h. Zn _____

i. Write overall cell reaction: $\text{Zn} + \text{Cu}^{2+} \rightarrow \text{Zn}^{2+} + \text{Cu}$ (10 marks) explain by writing arguments with redox reasoning for why it's a spontaneous reaction

Voltaic Cells Study Guide

Puja Mehta



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