

3.7) From A. S. Campbell and E. M. Kartzmark (The electrical conductance of strong electrolytes: A test of Stokes equation) *Can. J. Chem.* **7** (1955).

physical sciences data 41

Inorganic Electrolyte Solutions of AgNO₃ Solutions at 25, 35 and 95 °C

c	Λ _{exp}	Λ _{calc}	
		Δ = 2.3	Δ = 4.35
Handbook of Electrolyte Solutions			
I. Silver nitrate at 25% Λ _∞ = 133.36			
0.000	133.36	133.36	133.36
0.003	132.84	130.57	130.62
0.005	132.20	127.34	127.56
0.01	131.14	123.69	123.32
0.05	115.74	116.46	117.55
1	109.14	109.84	112.81
1.001	77.82	78.69	89.20
1.992	64.20	63.90	76.29
3.028	55.97	53.24	63.63
4.060	48.50	44.69	56.46
5.029	43.14	37.77	48.64
6.786	38.55	31.87	41.58
7.1	34.70	26.79	36.36
8.011	31.20	22.46	30.09
9.010	27.99	18.31	24.44
9.709	26.1	16.09	23.67
II. Silver nitrate at 35% Λ _∞ = 158.43			
0.986	152.3	129.7	133.2
1.4023	85.37	83.0	97.1
1.7757	79.55	76.8	91.6
3.0322	65.61	60.5	75.6
3.8230	58.91		66.8
4.2306	50.06		62.9
III. Silver nitrate at 95% Λ _∞ = 340.67			
0.0504	298.3		295.2
1	278.0		281.4
1.230	172.7		199.9
2.189	147.8		165.8
2.967	130.8		144.2
4.829	103.1		104.1
6.591	84.88		77.57
8.830	68.12		64.27
9.906	64.28		45.60
11	50.33		39.0

* Calculated using the temperature coefficients of conductance taken from Campbell's papers.

Handbook Of Electrolyte Solutions

ML Morrison



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