

- vii. **Name two non-metals which are both brittle and non-ductile**
Ans. Graphite and iodine are two non-metals which are brittle and non-ductile in nature.
- viii. **Name the most abundant non-metal in the earth's crust**
Ans. The most abundant non-metal in the earth's crust is oxygen. It is 47 % of earth's crust.
- ix. **Give the non-metallic trend in halogens**
Ans. The non-metallic trend in halogens decreases from top to bottom because of increasing atomic sizes.
- x. **Why do the non-metals accept electrons readily?**
Ans. Non-metals accept electrons readily because they are more electronegative and are usually electron deficient in nature. So they form anions by gaining electrons.



- xi. **Why non-metals do not react with dilute acids while metals do react?**
Ans. Non-metals do not react with dilute acids because non-metals are itself electron acceptors while metals react readily because they lose electrons readily.
- xii. **How can we distinguish a metal from a non-metal by simple physical methods?**
Ans.

Metals	Non-metals
i. Metals are good conductor of heat and electricity. ii. Metals possess high melting and boiling points. iii. They are lustrous.	i. Non metals are bad conductor of heat and electricity. ii. Non metals, possess low melting and boiling points. iii. They are dull.

- xiii. **How we can distinguish a substance is metal or non-metals with the help of an acid?**
Ans.

Metal	Non metal
Metals react with dilute acids easily because metals are itself electron donors $\text{Zn} + 2\text{HCl} \longrightarrow \text{ZnCl}_2 + \text{H}_2$ $\text{Zn} + \text{H}_2\text{SO}_4 \longrightarrow \text{ZnSO}_4 + \text{H}_2$	Non-metals or usually do not react with dilute acids because non-metals are itself electron accepters.

- xiv. **Why is HF a weak acid?**
Ans. HF is a weak acid because it does not release its proton easily due to presence of strong hydrogen bonding and it ionizes to a small extent in aqueous solution.
Fig: H- bonding in HF molecules

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