

Acids & Bases HOMEWORK

Lesson 1

1. Provide the name for each of the strong acids listed below. Provide the formula for each of the strong bases listed below.

- HCl ~~hydrochloric acid~~
- H₂SO₄ ~~hydrochloric acid~~
- HNO₃ ~~nitric acid~~
- HClO₄ ~~perchloric acid~~
- Strontium hydroxide ~~LiOH~~
- potassium hydroxide ~~KOH~~
- calcium hydroxide ~~CaOH~~
- calcium hydroxide ~~Ca(OH)₂~~
- ammonium hydroxide ~~Li(OH)₄~~

HClO₄ is STRONG with a lower cell completely dissociate in water

2. Of the acids and bases listed below, circle the ones that completely dissociate in water.

HCl

HF

H₂S

Mg(OH)₂

HNO₃

Ca(OH)₂

H₂SO₄

HClO₄

3. What is the formula for hydrosulfuric acid? **H₂S**

4. H₂C₂O₄ will react with water and lose a H⁺ ion in the process.

- Write a balanced equation for the reaction. **H₂C₂O₄ + H₂O → H₃C₂O₄⁻ + H₃O⁺**
- Would you expect a large or small equilibrium constant for this reaction? **small, because H₂C₂O₄ is a weak acid.**
- Would you define H₂C₂O₄ as a proton donor or a proton acceptor? **proton donor, because it is losing a H⁺ ion.**
- In your written equation, is H₂C₂O₄ acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid, because it is losing a proton.**
- Identify the conjugate base of H₂C₂O₄. **HC₂O₄⁻**
- Identify H₂C₂O₄ as polyprotic, amphoteric, or neither. **polyprotic, because it contains multiple H's to give.**

5. F⁻ will react with water and gain a H⁺ ion in the process.

- Write a balanced equation for the reaction. **F⁻ + H₂O → HF + OH⁻**
- Would you expect a large or small equilibrium constant for this reaction? **small, because F⁻ is a weak base.**
- Would you define F⁻ as a proton donor or a proton acceptor? **proton acceptor, because it is gaining a H⁺ ion.**
- In your written equation, is F⁻ acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry base, because it is gaining a proton.**
- Identify the conjugate acid of F⁻. **HF**
- Identify F⁻ as polyprotic, amphoteric, or neither. **neither**

6. HSO₄⁻ will react with water and lose a H⁺ ion in the process.

- Write a balanced equation for the reaction. **HSO₄⁻ + H₂O → SO₄²⁻ + H₃O⁺**
- Would you expect a large or small equilibrium constant for this reaction? **small, because HSO₄⁻ is a weak acid.**
- Would you define HSO₄⁻ as a proton donor or a proton acceptor? **proton donor, because it is losing a H⁺ ion.**
- In your written equation, is HSO₄⁻ acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid, because it is losing a proton.**
- Identify the conjugate base of HSO₄⁻. **SO₄²⁻**
- Identify HSO₄⁻ as polyprotic, amphoteric, or neither. **amphoteric, because of both a negative charge (only will take a proton) and a H to give (can lose the proton).**

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