

# 14.1

## HYDROGEN IONS AND ACIDITY

### Section Overview

#### Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and of aqueous concentrations.
2. Calculate hydroxide and hydronium ion concentrations and pH values and vice versa for weak acids and bases.
3. Calculate the pH of weak acid solutions.

#### Vocabulary

- |                   |                                    |                   |
|-------------------|------------------------------------|-------------------|
| 1. autoionization | 2. conjugate acid-base pair        | 3. weak acid      |
| 4. hydronium      | 5. hydroxide ion ( $\text{OH}^-$ ) | 6. autoionization |
|                   | 7. weak base                       | 8. pH             |

#### Key Equations

- |   |   |
|---|---|
| 1. $\text{pH} = -\log[\text{H}_3\text{O}^+]$ (Eq. 14.1) | 2. $\text{pOH} = -\log[\text{OH}^-]$ (Eq. 14.2) |
| 3. $\text{pH} + \text{pOH} = \text{pK}_w$ (Eq. 14.3)    | 4. $\text{pH} + \text{pOH} = 14.00$ (Eq. 14.4)  |

#### Part A: Completion

Use the appropriate number to identify the relationship of the conjugate acid-base pair indicated in the second step of the acid-base reaction given in the following examples.

1.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$  (water autoionization)
2.  $\text{H}_2\text{O} + \text{HCl} \rightleftharpoons \text{H}_3\text{O}^+ + \text{Cl}^-$  (hydrochloric acid ionization)
3.  $\text{H}_2\text{O} + \text{NH}_3 \rightleftharpoons \text{OH}^- + \text{NH}_4^+$  (ammonia ionization)
4.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$  (water autoionization)
5.  $\text{H}_2\text{O} + \text{HCl} \rightleftharpoons \text{H}_3\text{O}^+ + \text{Cl}^-$  (hydrochloric acid ionization)
6.  $\text{H}_2\text{O} + \text{NH}_3 \rightleftharpoons \text{OH}^- + \text{NH}_4^+$  (ammonia ionization)
7.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$  (water autoionization)
8.  $\text{H}_2\text{O} + \text{HCl} \rightleftharpoons \text{H}_3\text{O}^+ + \text{Cl}^-$  (hydrochloric acid ionization)
9.  $\text{H}_2\text{O} + \text{NH}_3 \rightleftharpoons \text{OH}^- + \text{NH}_4^+$  (ammonia ionization)

- |     |                |
|-----|----------------|
| 1.  | conjugate acid |
| 2.  | conjugate base |
| 3.  | acid           |
| 4.  | conjugate acid |
| 5.  | base           |
| 6.  | conjugate base |
| 7.  | acid           |
| 8.  | conjugate acid |
| 9.  | base           |
| 10. | conjugate base |
| 11. | conjugate acid |
| 12. | conjugate base |

# 19 2 Hydrogen Ions And Acidity Work Sheet Answers

**RC Schank**



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