

Unit 8 Review

1. The balanced equation  $\text{P}_4(\text{s}) + 6\text{H}_2(\text{g}) \rightarrow 4\text{PH}_3(\text{g})$  tells us that 4.0 mol  $\text{H}_2$   
 A) reacts with 2.0 mol  $\text{P}_4$   
 B) produces 8.0 mol  $\text{PH}_3$   
 C) cannot react with phosphorus  
 D) produces 2.7 mol  $\text{PH}_3$   
 E) reacts with 4.0 mol  $\text{P}_4$

2. A 1.5-mol sample of  $\text{Fe}_2\text{O}_3$  was decomposed according to the equation:  
 $2\text{Fe}_2\text{O}_3(\text{s}) \rightarrow 2\text{Fe}_2\text{O}(\text{s}) + 3\text{O}_2(\text{g})$

How many moles of  $\text{O}_2$  are formed assuming 100% yield?  
 A) 0.9 mol  
 B) 1.5 mol  
 C) 1.2 mol  
 D) 0.7 mol  
 E) 2.0 mol

3. The rusting of iron is represented by the equation  $4\text{Fe} + 3\text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3$ . If you have a 1.04-mol sample of iron, how many moles of  $\text{Fe}_2\text{O}_3$  will there be after the iron has rusted completely?

- A) 0.527 mol
- B) 0.790 mol
- C) 1.05 mol
- D) 1.58 mol
- E) 2.07 mol

4. Consider the following reaction, where X represents an unknown element.  
 $8\text{X}_2(\text{s}) + 2\text{B}_2\text{O}_3(\text{s}) \rightarrow \text{B}_2\text{X}_4(\text{s}) + 3\text{X}_2\text{O}_3(\text{g})$

If 175 g of X reacts completely with diboron trioxide to produce 1.43 mol of  $\text{B}_2\text{X}_4$ , what is the identity of X?

- A) Ge
- B) Mg
- C) Si
- D) Sn
- E) C

5. Refer to the following equation:  
 $4\text{NH}_3(\text{g}) + 7\text{O}_2(\text{g}) \rightarrow 4\text{NO}(\text{g}) + 6\text{H}_2\text{O}(\text{g})$

How many molecules of  $\text{NO}$  are produced when 3.02 mol of ammonia is completely reacted?

- A) 12.08
- B)  $3.64 \times 10^{23}$
- C)  $1.82 \times 10^{23}$
- D) 1.81
- E) none of these

6. In the reaction  $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightarrow 2\text{NH}_3(\text{g})$ , how many moles of ammonia would be produced from 1.09 mol of hydrogen and excess nitrogen?

- A) 1.42 mol
- B) 3.27 mol
- C) 0.727 mol
- D) 2.18 mol
- E) 0.363 mol

7. Refer to the following unbalanced equation:  
 $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$   
 What mass of oxygen ( $\text{O}_2$ ) is required to react completely with 13.1 g of  $\text{C}_2\text{H}_6$ ?

- A)  $2.99 \times 10^2$  g
- B) 23.1 g
- C) 4.86 g
- D) 49.2 g
- E) 0.192 g

8. How many molecules of carbon dioxide will be formed if 7.07 g of propane is burned in the following reaction?



- A)  $1.37 \times 10^{23}$  molecules
- B)  $1.03 \times 10^{23}$  molecules
- C)  $4.14 \times 10^{23}$  molecules
- D)  $3.10 \times 10^{23}$  molecules
- E)  $5.17 \times 10^{23}$  molecules

9. Calculate the mass of carbon dioxide produced from 15.7 g of octane,  $\text{C}_8\text{H}_{18}$ , in the following reaction.



- A) 79.2 g  $\text{CO}_2$
- B) 158 g  $\text{CO}_2$
- C) 89.1 g  $\text{CO}_2$
- D) 1.258 g  $\text{CO}_2$
- E) 1.400 g  $\text{CO}_2$

10. Calculate the molecules of oxygen required to react with 18.2 g of sulfur in the following reaction.



- A)  $1.03 \times 10^{23}$  molecules  $\text{O}_2$
- B)  $2.06 \times 10^{23}$  molecules  $\text{O}_2$
- C)  $1.71 \times 10^{23}$  molecules  $\text{O}_2$
- D)  $3.42 \times 10^{23}$  molecules  $\text{O}_2$
- E)  $5.13 \times 10^{23}$  molecules  $\text{O}_2$

# Multiple Choice Stoichiometry Test

**Michael Rajnik**



**Multiple Choice Stoichiometry Test:**

## Whispering the Strategies of Language: An Mental Journey through **Multiple Choice Stoichiometry Test**

In a digitally-driven world where screens reign great and instant connection drowns out the subtleties of language, the profound strategies and emotional subtleties hidden within words usually move unheard. However, nestled within the pages of **Multiple Choice Stoichiometry Test** a charming fictional prize pulsating with natural emotions, lies a fantastic journey waiting to be undertaken. Published by an experienced wordsmith, that wonderful opus encourages viewers on an introspective trip, softly unraveling the veiled truths and profound influence resonating within ab muscles cloth of each and every word. Within the psychological depths of this emotional evaluation, we will embark upon a heartfelt exploration of the book is primary styles, dissect its interesting publishing fashion, and yield to the strong resonance it evokes serious within the recesses of readers hearts.

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