

# Chemistry Study Guide

Periodic Table of the Elements

|    |    |    |    |    |    |    |    |    |    |    |    |     |    |     |    |     |     |   |    |    |    |    |    |    |    |    |    |
|----|----|----|----|----|----|----|----|----|----|----|----|-----|----|-----|----|-----|-----|---|----|----|----|----|----|----|----|----|----|
| 1  | 2  |    |    |    |    |    |    |    |    |    |    | 3   | 4  | 5   | 6  | 7   | 8   | 9 | 10 | 11 | 12 | 13 | 14 | 15 | 16 | 17 | 18 |
| H  | He |    |    |    |    |    |    |    |    |    |    | B   | C  | N   | O  | F   | Ne  |   |    |    |    |    |    |    |    |    |    |
| Li | Be |    |    |    |    |    |    |    |    |    |    | Al  | Si | P   | S  | Cl  | Ar  |   |    |    |    |    |    |    |    |    |    |
| Na | Mg | Sc | Ti | V  | Cr | Mn | Fe | Co | Ni | Cu | Zn | Ga  | Ge | As  | Se | Br  | Kr  |   |    |    |    |    |    |    |    |    |    |
| Rb | Sr | Y  | Zr | Nb | Mo | Tc | Ru | Rh | Pd | Ag | Cd | In  | Sn | Sb  | Te | I   | Xe  |   |    |    |    |    |    |    |    |    |    |
| Cs | Ba |    | Hf | Ta | W  | Re | Os | Ir | Pt | Au | Hg | Tl  | Pb | Bi  | Po | At  | Rn  |   |    |    |    |    |    |    |    |    |    |
| Fr | Ra |    | Rf | Db | Sg | Bh | Hs | Mt | Ds | Rg | Cn | Uut | Fl | Uup | Lv | Uus | Uuo |   |    |    |    |    |    |    |    |    |    |
|    |    | La | Ce | Pr | Nd | Pm | Sm | Eu | Gd | Tb | Dy | Ho  | Er | Tm  | Yb | Lu  |     |   |    |    |    |    |    |    |    |    |    |
|    |    | Ac | Th | Pa | U  | Np | Pu | Am | Cm | Bk | Cf | Es  | Fm | Md  | No | Lr  |     |   |    |    |    |    |    |    |    |    |    |

## Objective:

To understand the *structure of an atom*, and the *nature + properties* of subatomic particles 🎯

## Matter:

Anything that has **mass** and **occupies space** 🌟🌍

Examples: air, gases, liquids, solidz!

## Atoms:

Atoms are the **smallest, indivisible particle** 🔬👉

- **Outer part** → *shell/orbit/energy level* → carries the electrons ( $e^-$ )
- **Inner part** → *nucleus* → houses the protons ( $p^+$ ) and neutrons ( $n$ )
- Outer & Inner part → together these are the **subatomic particles**

## Subatomic Particle Table:

| Particle | Symbol | Mass (amu)       | Charge | Location    |
|----------|--------|------------------|--------|-------------|
| Proton   | $p^+$  | 1                | +1     | Nucleus     |
| Neutron  | $n$    | 1                | 0      | Nucleus     |
| Electron | $e^-$  | $\frac{1}{1836}$ | -1     | Orbit/Shell |

# Section 2 nuclear Chemistry Study Guide Key

**SJ Ball**



**Section 2 nuclear Chemistry Study Guide Key:**

## Reviewing **Section 2nuclear Chemistry Study Guide Key**: Unlocking the Spellbinding Force of Linguistics

In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is truly astonishing. Within the pages of "**Section 2nuclear Chemistry Study Guide Key**," an enthralling opus penned by a very acclaimed wordsmith, readers attempt an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve into the book's central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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