

## EXERCISE-1

1-a) Write a JAVA program to display default value of all primitive data types of JAVA  
class DefaultValues

```
{
    byte b;
    short s;
    int i;
    long l;
    float f;
    double d;
    char c;
    boolean bl;
    String s;
    public static void main(String[] args)
    {
        DefaultValues df=new DefaultValues();
        System.out.println("Byte :"+df.b);
        System.out.println("Short :"+df.s);
        System.out.println("Int :"+df.i);
        System.out.println("Long :"+df.l);
        System.out.println("Float :"+df.f);
        System.out.println("Double :"+df.d);
        System.out.println("Char :"+df.c);
        System.out.println("Boolean :"+df.bl);
        System.out.println("String :"+df.s);
    }
}
```

output:

```
javac DefaultValues.java
java DefaultValues
Byte :0
Short :0
Int :0
Long :0
Float :0.0
Double :0.0
Char :
Boolean :false
String: null
```

# Network Programming In Java Lab Manual

**Ensheng Dong**



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## **Table of Contents Network Programming In Java Lab Manual**

1. Understanding the eBook Network Programming In Java Lab Manual
  - The Rise of Digital Reading Network Programming In Java Lab Manual
  - Advantages of eBooks Over Traditional Books
2. Identifying Network Programming In Java Lab Manual
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Network Programming In Java Lab Manual
  - User-Friendly Interface
4. Exploring eBook Recommendations from Network Programming In Java Lab Manual
  - Personalized Recommendations
  - Network Programming In Java Lab Manual User Reviews and Ratings
  - Network Programming In Java Lab Manual and Bestseller Lists
5. Accessing Network Programming In Java Lab Manual Free and Paid eBooks
  - Network Programming In Java Lab Manual Public Domain eBooks
  - Network Programming In Java Lab Manual eBook Subscription Services
  - Network Programming In Java Lab Manual Budget-Friendly Options

6. Navigating Network Programming In Java Lab Manual eBook Formats
  - ePub, PDF, MOBI, and More
  - Network Programming In Java Lab Manual Compatibility with Devices
  - Network Programming In Java Lab Manual Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Network Programming In Java Lab Manual
  - Highlighting and Note-Taking Network Programming In Java Lab Manual
  - Interactive Elements Network Programming In Java Lab Manual
8. Staying Engaged with Network Programming In Java Lab Manual
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Network Programming In Java Lab Manual
9. Balancing eBooks and Physical Books Network Programming In Java Lab Manual
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Network Programming In Java Lab Manual
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Network Programming In Java Lab Manual
  - Setting Reading Goals Network Programming In Java Lab Manual
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Network Programming In Java Lab Manual
  - Fact-Checking eBook Content of Network Programming In Java Lab Manual
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

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for the nitration 4 01 ml of nitric acid 4 00 ml of sulfuric acid and 4 56 ml bromobenzene were used to complete the reaction with these measurements bromobenzene was found to be the limiting reagent since only 4 56 ml of bromobenzene was used the amount of product predicted to be recovered was 8 7533 g

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the aim of the experiment is to determine and study electrophilic aromatic substitution reaction eas particularly nitration of methyl benzoate go give methyl m nitrobenzoate thus determine the weight melting point and percentage yield of the pure methyl m nitrobenzoate abstract

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in this experiment we will carry out the nitration of bromobenzene bromine substituted on an aromatic ring is an ortho para directing group it will direct an incoming electrophile to the ortho and para positions it is also deactivating meaning that its reactions will be slower than the identical reactions of benzene

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section 20711 lab experiment nitration of bromobenzene introduction the chemical reaction of nitration consists of a nitro group being added to or substituted in a molecule nitration can basically be carried out by a mixture of concentrated nitric acid and sulfuric acid this mixture is useful to obtain the active nitronium ion

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1 go to the restroom and wash the affected area with lots of soap and water the soap in the lab is not suitable for this purpose 2 return to the lab and apply mineral oil to the affected area 3 the summary to this warning is not to touch any part of your body while performing this experiment if you must leave the lab for any reason 1

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in the nitration of bromobenzene a mixture of sulfuric acid and nitric acid react with one another nitric acid the weaker of the two acids will serve as base to form nitronium ion bromobenzene will attack the nitronium ion by nucleophilic attack

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then actual is your actual findings yield actual theoretical  $\times 100$  what are the possible side reactions of nitration of bromobenzene dinitration can occur at ortho para or at both ortho positions occurs in overheating meta product happens very rarely because bromine is an ortho para director

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the soap in the lab is not suitable for this purpose 2 return to the lab and apply mineral oil to the affected area theory and explanation of the nitration of bromobenzene if the temperature of the reaction mixture exceeds 60 °C

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the bromine substituent in bromobenzene is an ortho para director the para position is favored due to steric effects in the reaction that takes place in this lab a nitro group is substituted onto bromobenzene at the para position to form 1-bromo-4-nitrobenzene

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study with quizlet and memorize flashcards containing terms like reacting sulfuric and nitric acid together produces the nitronium ion alone nitric acid is not a strong enough electrophile to react with bromobenzene sulfuric acid is added to the nitric acid the reverse rxn can overheat and boil easily includes a claisen adaptor and

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over 5 minutes slowly pipet this bromobenzene into the icy cold nitration mixture  $\text{HNO}_3$   $\text{H}_2\text{SO}_4$  in the 125 ml flask and stir continue to stir the reaction mixture in the ice bath for 30 minutes do not heat in a new 250 ml erlenmeyer

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on the right the arenium cation has a nitro set with incredibly small 0.6 Å bond distances this will result in an optimization failure as the atoms are even close for the programme to accommodate their electrons demonstration of creating the p arenium cation by nitration of bromobenzene

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in this experiment the electrophilic aromatic substitution mechanism was used to obtain 4 nitrobromobenzene and 2 nitrobromobenzene recrystallization was used to purify 4 nitrobromobenzene

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the purpose of this lab was to observe where substituents are added on a benzene ring by way of electrophilic aromatic substitution the electrophilic source in the nitration of bromobenzene comes from a nitronium ion  $\text{NO}_2^+$  which acts as the electrophilic species

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nitration of bromobenzene is an electrophilic aromatic substitution reaction the electrophile will be formed from 4 0ml of nitric acid and 4 0ml of sulfuric acid and will then react with 4 5ml of bromobenzene this will produce o bromonitrobenzene and p bromonitrobenzene

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experimental spectra reaction a fluorobenzene stock fluorobenzene 1 h 1 h 19 f 19 f 13 c nmr fid for reference only sample 1 h 1 h 19 f 19 f 13 c nmr eas nitration of fluorobenzene crude product not available for submission for credit

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chem 2261 lab 5 nitration of bromobenzene introduction an electrophilic aromatic substitution reaction is an important process that allows for intermediate chemical structures to be used as precursors to pharmaceuticals and other industrially produced chemicals by adding different functional groups into an aromatic ring

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lab report 10 nitration of bromobenzene raekwon filmore cm 244 section 40 march 27 2018 introduction for this experiment nitration of bromobenzene was the focus of the lab the benzene is an aromatic compound and when it reacts with wither a mixture of sulfuric acid or nitric acid creates what is known as a nitro group

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