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web log H⁺ = 10^{-4.05} = 8.91 × 10⁻⁵ mol dm⁻³ log pK_a = 3.75 - 4.05 = -0.30 = 10^{-0.30} = 0.50 and HCOOH = 0.1 so HCOO⁻ = 0.200 mol dm⁻³ amount vol x

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web only the better candidates could write successive ionisation equations in part c while part d about the pH of equimolar strong and weak acid solutions was poorly answered

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web a substance may be assigned to one of our four conceivable categories it may be an acid or a base but in addition it may be both an acid and a base or it may be neither an acid nor

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web section 2 pH of strong acids number of protons released monoprotic acid acid that releases one H⁺ ion per molecule e.g.

hcl hydrochloric acid hno₃ nitric acid ch

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