

17.1

HYDROGEN IONS AND ACIDITY

Section Overview

Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and of aqueous concentrations.
2. Calculate hydroxide and hydronium ion concentrations and pH, and calculate hydroxide ion concentration from pH.
3. Calculate the pH of weak acid solutions.

Vocabulary

- | | | |
|-------------------|--------------------------------|--------|
| 1. autoionization | 2. conjugate acid-base pair | 3. pH |
| 4. hydronium ion | 5. hydronium ion concentration | 6. pOH |
| | 7. weak acid | |

Key Equations

- | | |
|---------------------------------------------------------|-------------------------------------------------|
| 1. $\text{pH} = -\log[\text{H}_3\text{O}^+]$ (Eq. 17.1) | 2. $\text{pOH} = -\log[\text{OH}^-]$ (Eq. 17.2) |
| 3. $\text{pH} + \text{pOH} = \text{pK}_w$ (Eq. 17.3) | 4. $\text{pH} + \text{pOH} = 14.00$ (Eq. 17.4) |

Part A: Completion

Use the appropriate number to indicate the correctness of the response to each item. (0 = incorrect; 1 = correct; 2 = partially correct; 3 = correct; 4 = completely correct.)

1. When a substance is _____, it is a weak hydroxide base.
2. _____ is the concentration of H_3O^+ ions present in a solution.
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**Thirumalaisamy P.
Velavancorresponding**



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