

9.1

THE ARITHMETIC OF EQUATIONS

SECTION REVIEW

Objectives

- Calculate the amount of reactants required or product formed in a stoichiometric process.
- Interpret balanced chemical equations in terms of interacting moles, representative particles, masses, and gas volume at STP.

Key Terms

- stoichiometry

Part A Completion

Copy this completion exercise for class, your understanding of the concepts and terms is clear and reinforced in this section. Write clearly on the lines provided with a pencil, silver pen, or marker.

- The coefficients in a balanced chemical equation indicate _____ 1. _____ the relative number of _____ 2. _____ reactants and products. All stoichiometric calculations begin with a _____ 3. _____ and _____ 4. _____ are conserved in every reaction; moles, volume, and representative particles may not be.
- In solving stoichiometric problems, conversion factors relating moles of reactants to _____ 5. _____ of products are used.
- If you assume _____ 6. _____, the equation also tells you about the volume of gas.

Part B True-False

Circle the word *if* in the statements to indicate they are (C) correct or (I) incorrect. Write (C) or (I) on the lines.

- _____ 7. The coefficients in a balanced chemical equation can be used to determine the mass relating reactants to products.
- _____ 8. The coefficients in a balanced chemical equation tell the relative numbers of reactants and products, expressed in any suitable units of volume.
- _____ 9. To calculate the mass of a molecule in grams, you can use the molar mass and Avogadro's number.

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J Dewey



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