

14.1

HYDROGEN IONS AND ACIDITY

Section Overview

Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and of aqueous concentrations.
2. Calculate hydroxide and hydronium ion concentrations and pH values and vice versa for a given pH.
3. Calculate the pH of weak acid solutions.

Vocabulary

- | | | |
|-------------------|-----------------------------|-------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. acid solution |
| 4. hydroxide ion | 5. hydronium ion | 6. autoionization |
| | 7. weak solution | 8. pH |

Key Equations

- | | |
|---|---|
| 1. $\text{pH} = -\log[\text{H}_3\text{O}^+]$ (Eq. 14.1) | 2. $\text{pOH} = -\log[\text{OH}^-]$ (Eq. 14.2) |
| 3. $\text{pH} + \text{pOH} = \text{pK}_w$ (Eq. 14.3) | 4. $\text{pH} + \text{pOH} = 14.00$ (Eq. 14.4) |

Part A: Completion

Use the appropriate number to identify the relationship of the conjugate acid-base pair indicated in (a)–(d). Label each relationship as either acid-base conjugate or neither.

- Water and hydroxide ion. **2** Neither. Hydroxide ion is the conjugate base of water.
- Hydrochloric acid and hydroxide ion. **1** Acid-base conjugate.
- Hydrochloric acid and hydronium ion. **1** Acid-base conjugate.
- Hydroxide ion and hydronium ion. **2** Neither. They are not conjugate acid-base pairs.
- H_2O and H_3O^+ . **1** Neither. The acid base pair is water and hydroxide ion.
- Water and protonated hydroxymethanesulfonate. **1** Acid-base conjugate.
- H_2O and hydroxide ion. **2** Neither. Hydroxide ion is the conjugate base of water.

- neither
- acid-base conjugate
- acid-base conjugate
- neither

19 2 Hydrogen Ions And Acidity Answer Key

Judd E. Hollander



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