

### 18.4 Acid-Base Chemistry: Theory

#### Multiple-Choice Questions

1. Which combination of quantities are dimensional quantities? (correct)
- (A) pressure, concentration, pH, amount of substance  
 (B) volume, molarity, pH, amount of substance  
 (C) volume, molarity, pH, concentration of ions  
 (D) amount of substance, pH, concentration of ions, concentration of substance

(A)  $\text{pH} = 10.0$ ,  $\text{pH} = 10.00$

(B)  $10 \text{ g H}_2\text{O}$ ,  $1 \text{ mol H}_2\text{O}$ ,  $1 \text{ g H}_2\text{O}$ ,  $1 \text{ mol H}_2\text{O}$ ,  $10 \text{ g H}_2\text{O}$ ,  $1 \text{ mol H}_2\text{O}$   
 (C)  $1 \text{ mol H}_2\text{O}$ ,  $1 \text{ mol H}_2\text{O}$

2. An acid-base indicator has a  $\text{pK}_a$  of 4.5. In a solution of the indicator, the ratio of the concentration of the acid form to the concentration of the base form is 10:1. What is the pH of the solution? (correct)

(A)  $\text{pH} = 4.5$ ,  $\text{pH} = 4.5$ ,  $\text{pH} = 4.5$ ,  $\text{pH} = 4.5$

(B)  $10 \text{ g H}_2\text{O}$ ,  $1 \text{ mol H}_2\text{O}$ ,  $1 \text{ g H}_2\text{O}$ ,  $1 \text{ mol H}_2\text{O}$ ,  $10 \text{ g H}_2\text{O}$ ,  $1 \text{ mol H}_2\text{O}$

3. An acid-base indicator has a  $\text{pK}_a$  of 4.5. In a solution of the indicator, the ratio of the concentration of the acid form to the concentration of the base form is 10:1. What is the pH of the solution? (correct)

(A)  $10 \text{ g H}_2\text{O}$ ,  $1 \text{ mol H}_2\text{O}$ ,  $1 \text{ g H}_2\text{O}$ ,  $1 \text{ mol H}_2\text{O}$

4. Which of the following is a weak acid? (correct)



(B)  $10 \text{ g H}_2\text{O}$ ,  $1 \text{ mol H}_2\text{O}$ ,  $1 \text{ g H}_2\text{O}$ ,  $1 \text{ mol H}_2\text{O}$

5. A solution of an acid has a pH of 4.5. The concentration of the acid is 0.1 mol/L. What is the concentration of the conjugate base? (correct)



(B)  $10 \text{ g H}_2\text{O}$ ,  $1 \text{ mol H}_2\text{O}$ ,  $1 \text{ g H}_2\text{O}$ ,  $1 \text{ mol H}_2\text{O}$ ,  $10 \text{ g H}_2\text{O}$ ,  $1 \text{ mol H}_2\text{O}$

6. An acid-base indicator has a  $\text{pK}_a$  of 4.5. In a solution of the indicator, the ratio of the concentration of the acid form to the concentration of the base form is 10:1. What is the pH of the solution? (correct)



# Multiple Choice Stoichiometry Test

**Roger Slade, Maureen Slade**



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