

Name Key

1. Hydrogen, H<sub>2</sub>, is used as rocket fuel. The hydrogen is burned in oxygen, O<sub>2</sub>, according to the following equation. What is the enthalpy change if 200.0 g of hydrogen, H<sub>2</sub> is burned?



$$\frac{200.0 \text{ g H}_2}{2.016 \text{ g H}_2} \times \frac{1 \text{ mol H}_2}{1 \text{ mol H}_2} = 99.2 \text{ mol H}_2$$

$$99.2 \text{ mol H}_2 \times \frac{-484 \text{ kJ}}{2 \text{ mol H}_2} = -24,006 \text{ kJ}$$

-24,006 kJ  
~~24,006 kJ~~

2. Carbon disulfide burns in air, producing carbon dioxide and sulfur dioxide.



What is  $\Delta H$  for the following reaction?

reverse so  
change sign of  
 $\Delta H$  = out in 1/2  
quantity



3. Following is the equation for the combustion of acetone:



How much heat will be given off (in kcal) if 4.5 mol of acetone is burned completely?

$$\frac{-853.6 \text{ kcal}}{2 \text{ mol C}_3\text{H}_6\text{O}} \times 4.5 \text{ mol C}_3\text{H}_6\text{O} = -1920.6 \text{ kcal}$$

Exothermic

4. The specific heat of iron is 0.110 cal / g · °C. How much heat must be added to 5.2 g of iron at 25 °C to raise its temperature to 225 °C?

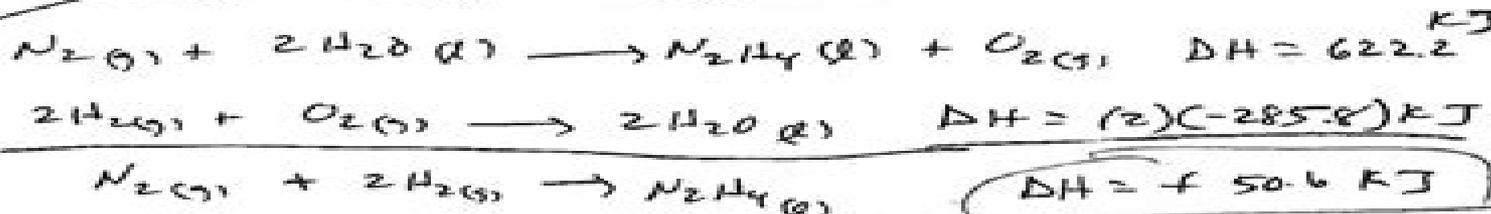
$$\text{heat} = q = m c \Delta T = (5.2 \text{ g}) \left( \frac{0.110 \text{ cal}}{\text{g} \cdot \text{C}} \right) (225 - 25 \text{ C}) = 114.4 \text{ cal}$$

Endothermic

5. Hydrazine, N<sub>2</sub>H<sub>4</sub>, is a colorless liquid used as a rocket fuel. What is the enthalpy change for the process in which hydrazine is formed from its elements (reaction (1) below)? Use Hess's Law to calculate  $\Delta H$  for reaction (1) given the data below.



Reverse  
 double quantity



# 19 3 Chemistry Practice Problems Answer Key

**Ensheng Dong**



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