

14.1

HYDROGEN IONS AND ACIDITY

Section Overview

Objectives

1. Identify conjugate acid-base pairs in these pairs of hydrolyzed acid-base species, and write the corresponding chemical equations.
2. Calculate K_a and K_b for weak acids and bases, and K_a and K_b for the conjugate acid-base pairs.
3. Calculate pH and pOH for acids and bases.

Vocabulary

- | | | |
|-------------------|-----------------------------|------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. acid solution |
| 4. autoionization | 5. K_a | 6. K_b |
| | 7. pH | |

Key Equations

- | | |
|--|--|
| 1. $\text{pH} = -\log[\text{H}^+]$ (Equation 14.1) | 2. $\text{pOH} = -\log[\text{OH}^-]$ (Equation 14.2) |
| 3. $\text{pH} + \text{pOH} = \text{pK}_w$ | 4. $\text{pH} + \text{pOH} = 14.00$ |

Part A: Completion

Use the appropriate number of significant figures in the answers to the following questions that are based on the following data: K_a for acetic acid is 1.8×10^{-5} and K_b for ammonia is 1.8×10^{-5} .

1. What products are **1** formed when hydrochloric acid (HCl) reacts with H_2O ? The concentration of these acids produced in 0.100 M HCl solution is **2** mol/L .
2. Hydrochloric acid dissociates from **3** HCl to H^+ and Cl^- ions. **4** HCl is a strong acid. The K_a of HCl is **5** 1.0×10^7 . **6** HCl is a strong acid. **7** HCl is a strong acid. **8** HCl is a strong acid.
3. **9** HCl is a strong acid. **10** HCl is a strong acid. **11** HCl is a strong acid. **12** HCl is a strong acid. **13** HCl is a strong acid. **14** HCl is a strong acid. **15** HCl is a strong acid. **16** HCl is a strong acid. **17** HCl is a strong acid. **18** HCl is a strong acid. **19** HCl is a strong acid. **20** HCl is a strong acid.

21. HCl
22. 1.0×10^7
23. 1.0×10^7
24. 1.0×10^7
25. 1.0×10^7
26. 1.0×10^7
27. 1.0×10^7
28. 1.0×10^7
29. 1.0×10^7
30. 1.0×10^7

19 2 Hydrogen Ions And Acidity Answers

**John L. Tymoczko, Jeremy M.
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