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GETTING SMARTER

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# Getting Smarter

**M Mosston**



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web solution moles of  $CH_3OH$  5.532017 mole moles of  $H_2O$  40.1822 moles therefore according to the equation mole fraction of  $CH_3OH$   $\frac{5.532017}{5.532017 + 40.1822}$  0.12073 to solve more examples on mole fraction formula and practice more question please visit byju s com

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web you may work out the mole fractions on your own problem 5 what is the mass of a sample of a 0.449 molal kbr that contains 2.92 kg of water 1 molality moles solute divided by kilograms solute 0.449 mol kg<sup>-1</sup> x 2.92 kg x 131.108 mol of kbr 2 moles times molar mass equals grams 131.108 mol 119.0023 g mol 156 g kbr

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web calculate the mass percent mole fraction molality and molarity of hexane  $C_6H_{14}$  0.660 g cm<sup>-3</sup> when it is mixed with heptane  $C_7H_{16}$  0.684 g cm<sup>-3</sup> in a 35.0550 ml ratio assume that the volumes are additive

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web notice that the mole fraction has no units on it and is written as a decimal value do not change it to percent note of caution you could see the term mole percent it is simply the mole fraction multiplied by 100 for example in the problem just below the mole fraction of cinnamic acid is 0.2885 its mole percent would be 28.85

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web since the mole fraction of the o<sub>2</sub> is 0.050 it contributes 0.050 or if you will 5.0 of the total pressure thus total pressure is 0.210 0.050 4.2 atm problem 9 a sample of 1.43 g of helium and an unweighed quantity of o<sub>2</sub> are mixed in a flask at room temperature

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