

18.4 Molecular Structure: Mass

Multiple Choice:

1. The molar masses of quantities in chemical equations are based on _____
 (A) percent composition (B) percent yield
 (C) molar mass (D) molar volume
2. Percent of oxygen in $\text{CaSO}_4 \cdot \frac{1}{2}\text{H}_2\text{O}$ is approximately (to two significant figures):

(A) 20% (B) 25% (C) 30% (D) 35%

3. In a gas the other reactions represent copper metal and silver nitrate, 2.5 g of Ag are precipitated by the reaction, and only 0.5 g of Ag⁺ ions remain in the solution.

What is the molar mass of the precipitate?

(A) 100 g/mol (B) 150 g/mol (C) 200 g/mol (D) 250 g/mol

4. In the reaction $\text{ZnCO}_3 \rightarrow \text{ZnO} + \text{CO}_2$, what is the mass of ZnO produced from 100 g of ZnCO_3 ?

(A) 65 g (B) 80 g (C) 90 g (D) 100 g

5. What is the molar mass of the gas produced in the following reaction?



(A) 44 g/mol (B) 46 g/mol (C) 48 g/mol (D) 50 g/mol

6. Aluminum reacts with H₂ gas to produce hydrogen gas. Calculate the mass of H₂ gas produced from 100 g of aluminum.



(A) 100 g (B) 150 g (C) 200 g (D) 300 g (E) 450 g

7. In a full combustion, structure-molecular formula $\text{C}_2\text{H}_5\text{CO}_2\text{H}$ what mass of CO_2 is produced from 100 g of $\text{C}_2\text{H}_5\text{CO}_2\text{H}$?



(A) 100 g (B) 150 g (C) 200 g (D) 300 g (E) 450 g

8. In a full combustion, structure-molecular formula $\text{C}_2\text{H}_5\text{CO}_2\text{H}$ what mass of CO_2 is produced from 100 g of $\text{C}_2\text{H}_5\text{CO}_2\text{H}$?



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