

Chemistry 2nd Semester Final Review **KEY**

This is meant as a guide. It does NOT contain every single thing you need to know for the final exam. Review notes, benchmark review sheets, homeworks, chapter study guides, etc.

Bring your book to school by Monday for turn in (seniors should bring book to bookstore by themselves ASAP)

Gas Laws (Ch. 13):

Subtopics: atmospheric pressure, Dalton's law of partial pressures, combined gas law, Ideal gas law, gas stoichiometry (22.4 L/mol for a gas at STP)

Practice questions:

- If a gas occupies 3.8 L at a pressure of 2.71 atm, what would the pressure be if the volume changes to 1.47 L? (Temp. and amt. of gas constant)
 $V_1=3.8 \text{ L}; P_1=2.71 \text{ atm}$
 $V_2=1.47 \text{ L}; P_2=? \text{ Atm}$
Since all units are standard (L, atm), we do not need to convert the units before proceeding.
 $(P_1V_1)/T_1=(P_2V_2)/T_2$
T goes away (not mentioned in problem)
 $P_1V_1=P_2V_2$
Isolate for P₂: $P_2=(P_1V_1)/V_2$
 $P_2=(2.71 \text{ atm} \times 3.8 \text{ L})/(1.47 \text{ L})=7.0 \text{ atm}$ (L cancelled and it's 2 sig figs)
- A gas has a volume of 1.49 L at a temperature of 34.75 degrees Celsius. What would the volume be at 78.41 degrees Celsius? (pressure & amt. of gas constant)
 $V_1=1.49 \text{ L}, T_1=34.75 \text{ degrees Celsius}$
 $V_2=? \text{ L}, T_2=78.41 \text{ degrees Celsius}$
We need to convert from Celsius to Kelvins before proceeding.
 $T_1=34.75 + 273=307.75 \text{ K}$
 $T_2=78.41 + 273=351.41 \text{ K}$
 $(P_1V_1)/T_1=(P_2V_2)/T_2$
Get rid of P
 $V_1/T_1=V_2/T_2$
Isolate V₂: $V_2=(V_1/T_1)T_2$
 $V_2=(1.49 \text{ L}/307.75 \text{ K})(351.41 \text{ K})=1.70 \text{ L}$
- What volume is occupied by 8.47 g of hydrogen gas at 84.7 degrees Celsius and 1.04 atm?
The combined gas law won't work here. You only have one condition and nothing changes.
 $PV=nRT$ will work. (n= moles and R= 0.08206 L atm/K mol)
 $P=1.04 \text{ atm}, V=? \text{ L}, n=?$ OOPS we need to convert from grams to moles!, $T=84.7 \text{ degrees Celsius} + 273=357.7 \text{ K}$
n: Hydrogen gas is diatomic!
 $8.47 \text{ g H}_2 \times (1 \text{ mol H}_2/2.02 \text{ g H}_2)=4.193069307 \text{ mol H}_2$
Isolate V: $V=(nRT)/P=(4.193069307 \text{ mol H}_2 \times 0.08206 \text{ L atm/K mol} \times 357.7 \text{ K})/1.04 \text{ atm}=118 \text{ L}$
- What volume is occupied by 56.75 g of oxygen gas at STP?

2nd Semester Final Review Guide Chemistry

CO Houle



2nd Semester Final Review Guide Chemistry:

This book delves into 2nd Semester Final Review Guide Chemistry. 2nd Semester Final Review Guide Chemistry is a vital topic that must be grasped by everyone, ranging from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into 2nd Semester Final Review Guide Chemistry, encompassing both the fundamentals and more intricate discussions.

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- Chapter 1: Introduction to 2nd Semester Final Review Guide Chemistry
- Chapter 2: Essential Elements of 2nd Semester Final Review Guide Chemistry
- Chapter 3: 2nd Semester Final Review Guide Chemistry in Everyday Life
- Chapter 4: 2nd Semester Final Review Guide Chemistry in Specific Contexts
- Chapter 5: Conclusion

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4. In chapter 3, this book will examine the practical applications of 2nd Semester Final Review Guide Chemistry in daily life. The third chapter will showcase real-world examples of how 2nd Semester Final Review Guide Chemistry can be effectively utilized in everyday scenarios.

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6. In chapter 5, this book will draw a conclusion about 2nd Semester Final Review Guide Chemistry. The final chapter will summarize the key points that have been discussed throughout the book.

This book is crafted in an easy-to-understand language and is complemented by engaging illustrations. It is highly recommended for anyone seeking to gain a comprehensive understanding of 2nd Semester Final Review Guide Chemistry.

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